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	AgNO₃ (Ag⁺¹)	Pb(NO₃)₂ (Pb⁺²)	Ca(NO₃)₂ Ca⁺²)
Na ₂ CO ₃ (CO ₃ ⁻²)			
Na ₂ S (S ⁻²)			
NaOH (OH ⁻¹)			
Na ₂ SO ₃ (SO ₃ ⁻²)			
NaCl (Cl ⁻¹)			

	AgNO_3 (Ag^{+1})	$\text{Pb}(\text{NO}_3)_2$ (Pb^{+2})	$\text{Ca}(\text{NO}_3)_2$ (Ca^{+2})
Na_2CO_3 (CO_3^{-2})	a) pink		
Na_2S (S^{-2})			
NaOH (OH^{-1})			
Na_2SO_3 (SO_3^{-2})			
NaCl (Cl^{-1})			

	AgNO_3 (Ag^{+1})	$\text{Pb}(\text{NO}_3)_2$ (Pb^{+2})	$\text{Ca}(\text{NO}_3)_2$ (Ca^{+2})
Na_2CO_3 (CO_3^{-2})	a) pink		
Na_2S (S^{-2})	b) Black		
NaOH (OH^{-1})			
Na_2SO_3 (SO_3^{-2})			
NaCl (Cl^{-1})			

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink		
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black		
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown		
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$			
$\text{NaCl} (\text{Cl}^{+1})$			

	AgNO_3 (Ag^{+1})	$\text{Pb}(\text{NO}_3)_2$ (Pb^{+2})	$\text{Ca}(\text{NO}_3)_2$ (Ca^{+2})
Na_2CO_3 (CO_3^{-2})	a) pink		
Na_2S (S^{-2})	b) Black		
NaOH (OH^{-1})	c) Dark pink./ brown		
Na_2SO_3 (SO_3^{-2})	d) NR		
NaCl (Cl^{-1})			

	AgNO_3 (Ag^{+1})	$\text{Pb}(\text{NO}_3)_2$ (Pb^{+2})	$\text{Ca}(\text{NO}_3)_2$ (Ca^{+2})
Na_2CO_3 (CO_3^{-2})	a) pink		
Na_2S (S^{-2})	b) Black		
NaOH (OH^{-1})	c) Dark pink./ brown		
Na_2SO_3 (SO_3^{-2})	d) NR		
NaCl (Cl^{-1})	e) white		

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black		
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown		
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR		
$\text{NaCl} (\text{Cl}^{-1})$	e) white		

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown		
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR		
$\text{NaCl} (\text{Cl}^{+1})$	e) white		

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown	h) NR	
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR		
$\text{NaCl} (\text{Cl}^{+1})$	e) white		

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown	h) NR	
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR	i) white	
$\text{NaCl} (\text{Cl}^{-1})$	e) white		

	AgNO_3 (Ag^{+1})	$\text{Pb}(\text{NO}_3)_2$ (Pb^{+2})	$\text{Ca}(\text{NO}_3)_2$ (Ca^{+2})
Na_2CO_3 (CO_3^{-2})	a) pink	f) white	
Na_2S (S^{-2})	b) Black	g) black	
NaOH (OH^{-1})	c) Dark pink./ brown	h) NR	
Na_2SO_3 (SO_3^{-2})	d) NR	i) white	
NaCl (Cl^{-1})	e) white	j) white	

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown	h) NR	
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR	i) white	
$\text{NaCl} (\text{Cl}^{-1})$	e) white	j) white	

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	k) white
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown	h) NR	
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR	i) white	
$\text{NaCl} (\text{Cl}^{+1})$	e) white	j) white	

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	k) white
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	l) NR
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown	h) NR	
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR	i) white	
$\text{NaCl} (\text{Cl}^{-1})$	e) white	j) white	

	AgNO_3 (Ag^{+1})	$\text{Pb}(\text{NO}_3)_2$ (Pb^{+2})	$\text{Ca}(\text{NO}_3)_2$ (Ca^{+2})
Na_2CO_3 (CO_3^{-2})	a) pink	f) white	k) white
Na_2S (S^{-2})	b) Black	g) black	l) NR
NaOH (OH^{-1})	c) Dark pink./ brown	h) NR	m) white
Na_2SO_3 (SO_3^{-2})	d) NR	i) white	
NaCl (Cl^{-1})	e) white	j) white	

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	k) white
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	l) NR
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown	h) NR	m) white
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR	i) white	n) NR
$\text{NaCl} (\text{Cl}^{-1})$	e) white	j) white	

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	k) white
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	l) NR
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown	h) NR	m) white
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR	i) white	n) NR
$\text{NaCl} (\text{Cl}^{-1})$	e) white	j) white	o) NR

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$	a) pink	f) white	k) white
$\text{Na}_2\text{S} (\text{S}^{-2})$	b) Black	g) black	l) NR
$\text{NaOH} (\text{OH}^{-1})$	c) Dark pink./ brown	h) NR	m) white
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$	d) NR	i) white	n) NR
$\text{NaCl} (\text{Cl}^{-1})$	e) white	j) white	o) NR

	$\text{AgNO}_3 (\text{Ag}^{+1})$	$\text{Pb}(\text{NO}_3)_2 (\text{Pb}^{+2})$	$\text{Ca}(\text{NO}_3)_2 (\text{Ca}^{+2})$
$\text{Na}_2\text{CO}_3 (\text{CO}_3^{-2})$			
$\text{Na}_2\text{S} (\text{S}^{-2})$			
$\text{NaOH} (\text{OH}^{-1})$			
$\text{Na}_2\text{SO}_3 (\text{SO}_3^{-2})$			
$\text{NaCl} (\text{Cl}^{+1})$			

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 +$
	$+ \text{AgNO}_3$	NO_3
	$+ \text{AgNO}_3$	NO_3
	$+ \text{AgNO}_3$	NO_3

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 + \text{CO}_3$ Anion Anion

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 + \text{CO}_3$ Anion Anion Charge: -1 -2

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 + \text{CO}_3$ Anion Anion Charge: -1 -2
	Na Ag Cation Cation Charge: +1 +1	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 + \text{CO}_3$ Anion Anion Charge: -1 -2
	Na Ag Cation Cation Charge: +1 +1	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 +$
	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
	IS THIS BALANCED:	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 +$
	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
	IS THIS BALANCED:	NO

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 +$
	$\text{Na}_2\text{CO}_3 + 2 \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
	Is This Balanced:	Yes

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 +$
	$\text{Na}_2\text{CO}_3 + 2 \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
	Is This Balanced:	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$\text{NO}_3 +$
	$\text{Na}_2\text{CO}_3 + 2 \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
	Is This Balanced:	Yes

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
	$+ \text{AgNO}_3$	NO_3
	$+ \text{AgNO}_3$	NO_3
	$+ \text{AgNO}_3$	NO_3

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	NO_3

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	NO_3

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
		Anion Anion

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation	Anion Anion
	Na Ag	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation	Anion Anion
	Na Ag Charge: +1 +1	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
		Anion Anion Charge: -1 -2

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
		Anion Anion Charge: -1 -2

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{S}$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{S}$
	IS THIS BALANCED?	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{S}$
	IS THIS BALANCED?	NO

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{S}$
	IS THIS BALANCED?	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{S}$
	IS THIS BALANCED?	NO

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2
b	$\text{Na}_2\text{S} + 2 \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{S}$
	IS THIS BALANCED?	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -2
b	$\text{Na}_2\text{S} + 2 \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{S}$
	IS THIS BALANCED?	YES

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	NO_3
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
c		Anion Charge: -1 Anion -2

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 +$
d	$+ \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 + \text{S}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 + \text{OH}$

Balanced Reactions

	Reactant	Product
a		
b		
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 + \text{OH}$

Balanced Reactions

	Reactant	Product
a		
b		
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 + \text{OH}$

Balanced Reactions

	Reactant	Product
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 + \text{OH}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 + \text{OH}$

Balanced Reactions

	Reactant	Product
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 + \text{OH}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 + \text{OH}$

Balanced Reactions

	Reactant	Product
c	NaOH + AgNO ₃	NO ₃ + OH
		Anion anion

Balanced Reactions

	Reactant	Product
c	NaOH + AgNO ₃	NO ₃ + OH
		Anion Anion
		Charge: -1 -1

Balanced Reactions

	Reactant	Product
c	NaOH + AgNO ₃	NO ₃ + OH
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -1

Balanced Reactions

	Reactant	Product
c	NaOH + AgNO ₃	NO ₃ + OH
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -1
	NaOH + AgNO ₃	NO ₃ + OH

Balanced Reactions

	Reactant	Product
c	NaOH + AgNO ₃	NO ₃ + OH
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -1
	NaOH + AgNO ₃	NaNO ₃ + AgOH

Balanced Reactions

	Reactant	Product
c	NaOH + AgNO ₃	NO ₃ + OH
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -1
	NaOH + AgNO ₃	NaNO ₃ + AgOH

Balanced Reactions

	Reactant	Product
c	NaOH + AgNO ₃	NO ₃ + OH
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -1
	NaOH + AgNO ₃	NaNO ₃ + AgOH
	Is This Balance?	

Balanced Reactions

	Reactant	Product
c	NaOH + AgNO ₃	NO ₃ + OH
	Cation Cation Na Ag Charge: +1 +1	Anion Anion Charge: -1 -1
	NaOH + AgNO ₃	NaNO ₃ + AgOH
	Is This Balance?	Yes

Balanced Reactions

	Reactant	Product
e	NaOH + AgNO ₃	NO ₃ + OH
		Anion Anion
		Charge: -1 -1
	Cation Cation Na Ag Charge: +1 +1	

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{S}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgOH}$
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{S}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgOH}$
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{S}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgOH}$
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{S}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgOH}$
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{S}$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgOH}$
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2
	Cation Cation	

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2
	Cation Cation Charge: +1 +1	

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2
	Cation Cation Charge: +1 +1	

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2
	Cation Cation Charge: +1 +1	
	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{SO}_4$

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2
	Cation Cation Charge: +1 +1	
	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{SO}_4$
	IS This BALANCED?	

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2
	Cation Cation Charge: +1 +1	
	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{SO}_4$
	IS This BALANCED?	NO

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2
	Cation Cation Charge: +1 +1	
	$\text{Na}_2\text{SO}_4 + 2 \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{SO}_4$
	IS This BALANCED?	Yes

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 + \text{SO}_4$
		anion anion Charge: -1 -2
	Cation Cation Charge: +1 +1	
	$\text{Na}_2\text{SO}_4 + 2 \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{SO}_4$
	IS This BALANCED?	Yes

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$
	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$
d	$\text{Na}_2\text{SO}_4 + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
a		
b		
c		
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgCl}$

Balanced Reactions

	Reactant	Product
a		
b		
c		
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgCl}$

Balanced Reactions

	Reactant	Product
a		
b		
c		
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NO}_3^- + \text{Cl}^-$ Anion Anion Charge: -1 -1

Balanced Reactions

	Reactant	Product
a		
b		
c		
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgCl}$
	Is This Balanced?	

Balanced Reactions

	Reactant	Product
a		
b		
c		
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NaNO}_3 + \text{AgCl}$
	Is This Balanced?	Yes

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 +$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 +$
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
a	$\text{Na}_2\text{CO}_3 + \text{AgNO}_3$	$2 \text{NaNO}_3 + \text{Ag}_2\text{CO}_3$
b	$\text{Na}_2\text{S} + \text{AgNO}_3$	$\text{NO}_3 +$
c	$\text{NaOH} + \text{AgNO}_3$	$\text{NO}_3 +$
e	$\text{NaCl} + \text{AgNO}_3$	$\text{NO}_3 +$

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{CO}_3$
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{S}$
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{OH}$
i	$\text{Na}_2\text{SO}_4 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{SO}_4$

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{CO}_3$
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{S}$
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{OH}$
i	$\text{Na}_2\text{SO}_4 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{SO}_4$

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{CO}_3$
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{S}$
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{OH}$
i	$\text{Na}_2\text{SO}_4 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{SO}_4$

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{CO}_3$
		Anion Anion

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{CO}_3$
		Anion Anion
		Charge: -3 -2

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{CO}_3$
		Anion Anion
		Charge: -3 -2

Balanced Reactions

	Reactant	Product						
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{CO}_3$						
		Anion Anion						
		Charge: -3 -2						
	<table style="margin-left: auto; margin-right: auto;"> <tr> <td>Cation</td> <td>Cation</td> </tr> <tr> <td>Na</td> <td>Pb</td> </tr> <tr> <td>Charge: +1</td> <td>+3</td> </tr> </table>	Cation	Cation	Na	Pb	Charge: +1	+3	
Cation	Cation							
Na	Pb							
Charge: +1	+3							

Balanced Reactions

	Reactant	Product						
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3) + \text{CO}_3$						
		Anion Anion						
		Charge: -1 -2						
	<table style="margin-left: auto; margin-right: auto;"> <tr> <td style="text-align: center;">Cation</td> <td style="text-align: center;">Cation</td> </tr> <tr> <td style="text-align: center;">Na</td> <td style="text-align: center;">Pb</td> </tr> <tr> <td style="text-align: center;">Charge: +1</td> <td style="text-align: center;">+2</td> </tr> </table>	Cation	Cation	Na	Pb	Charge: +1	+2	
Cation	Cation							
Na	Pb							
Charge: +1	+2							
	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_2$	$\text{NaNO}_3 + \text{Pb}_3(\text{CO}_3)_2$						
	Does this Balance?							

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3) + \text{CO}_3$
		<p>Anion Anion</p> <p>Charge: -1 -2</p>
	<p>Cation Cation</p> <p>Na Pb</p> <p>Charge: +1 +2</p>	
	$6 \text{Na}_2\text{CO}_3 + 6 \text{Pb}(\text{NO}_3)_2$	$12 \text{NaNO}_3 + 2 \text{Pb}_3(\text{CO}_3)_2$
	Does this Balance?	

Balanced Reactions

	Reactant	Product						
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3) + \text{CO}_3$						
		Anion Anion						
		Charge: -1 -2						
	<table style="margin-left: auto; margin-right: auto;"> <tr> <td>Cation</td> <td>Cation</td> </tr> <tr> <td>Na</td> <td>Pb</td> </tr> <tr> <td>Charge: +1</td> <td>+2</td> </tr> </table>	Cation	Cation	Na	Pb	Charge: +1	+2	
Cation	Cation							
Na	Pb							
Charge: +1	+2							
	$6 \text{Na}_2\text{CO}_3 + 6 \text{Pb}(\text{NO}_3)_2$	$12 \text{NaNO}_3 + 2 \text{Pb}_3(\text{CO}_3)_2$						
	Does this Balance?	NO						

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{CO}_3$
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{S}$

Balanced Reactions

	Reactant	Product
f		
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{S}$

Balanced Reactions

	Reactant	Product
f		
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{S}$ Anion Anion
		Charge: -1 -2

Balanced Reactions

	Reactant	Product
f		
g	Na_2S + $\text{Pb}(\text{NO}_3)_2$ Cation Cation	$(\text{NO}_3)_2$ + S Anion Anion
	Charge: +1 +2	Charge: -1 -2

Balanced Reactions

	Reactant	Product
f		
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Cation Cation	$(\text{NO}_3)_2 + \text{S}$ Anion Anion
	Charge: +1 +2	Charge: -1 -2
		$\text{NaNO}_3 + \text{PbS}$

Balanced Reactions

	Reactant	Product
f		
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Cation Cation	$(\text{NO}_3)_2 + \text{S}$ Anion Anion
	Charge: +1 +2	Charge: -1 -2
		$\text{NaNO}_3 + \text{PbS}$

Balanced Reactions

	Reactant	Product
f		
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Cation Cation	$(\text{NO}_3)_2 + \text{S}$ Anion Anion
	Charge: +1 +2	Charge: -1 -2
	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Is this balanced?	$\text{NaNO}_3 + \text{PbS}$

Balanced Reactions

	Reactant	Product
f		
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Cation Cation	$(\text{NO}_3)_2 + \text{S}$ Anion Anion
	Charge: +1 +2	Charge: -1 -2
	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Is this balanced?	$\text{NaNO}_3 + \text{PbS}$

Balanced Reactions

	Reactant	Product
f		
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Cation Cation	$(\text{NO}_3)_2 + \text{S}$ Anion Anion
	Charge: +1 +2	Charge: -1 -2
	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Is This Balanced?	$2 \text{NaNO}_3 + \text{PbS}$

Balanced Reactions

	Reactant	Product
f		
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Cation Cation	$(\text{NO}_3)_2 + \text{S}$ Anion Anion
	Charge: +1 +2	Charge: -1 -2
	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$ Is This Balanced?	$2 \text{NaNO}_3 + \text{PbS}$ YES

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{CO}_3$
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{S}$
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{OH}$

Balanced Reactions

	Reactant	Product
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{OH}$

Balanced Reactions

	Reactant	Product
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{OH}$
		Anion Charge: -1 Anion -1

Balanced Reactions

	Reactant	Product
h	NaOH + Pb(NO ₃) ₂	(NO ₃) ₂ + OH
	Cation Na	Cation Pb
		Anion Charge: -1
		Anion -1

Balanced Reactions

	Reactant	Product
h	NaOH + Pb(NO ₃) ₂	(NO ₃) ₂ + OH
	Cation Cation Na Pb Charge: +1 +2	Anion Anion Charge: -1 -1

Balanced Reactions

	Reactant	Product
h	NaOH + Pb(NO ₃) ₂	(NO ₃) ₂ + OH
	Cation Cation Na Pb Charge: +1 +2	Anion Anion Charge: -1 -1
h	NaOH + Pb(NO ₃) ₂	Pb(NO ₃) + NaOH

Balanced Reactions

	Reactant	Product
h	NaOH + Pb(NO ₃) ₂	(NO ₃) ₂ + OH
	Cation Cation Na Pb Charge: +1 +2	Anion Anion Charge: -1 -1
h	NaOH + Pb(NO ₃) ₂	Na(NO ₃) + Pb(OH)
	Is this Balanced?	NO

Balanced Reactions

	Reactant	Product
h	NaOH + Pb(NO ₃) ₂	(NO ₃) ₂ + OH
	Cation Cation Na Pb Charge: +1 +2	Anion Anion Charge: -1 -1
h	NaOH + Pb(NO ₃) ₂	NaNO ₃ + Pb(OH) ₂
	Is this Balanced?	No

Balanced Reactions

	Reactant	Product
h	NaOH + Pb(NO ₃) ₂	(NO ₃) ₂ + OH
	Cation Cation Na Pb Charge: +1 +2	Anion Anion Charge: -1 -1
h	2 NaOH + Pb(NO ₃) ₂	2 NaNO ₃ + Pb(OH) ₂
	Is this Balanced?	

Balanced Reactions

	Reactant	Product
h	NaOH + Pb(NO ₃) ₂	(NO ₃) ₂ + OH
	Cation Cation Na Pb Charge: +1 +2	Anion Anion Charge: -1 -1
h	2 NaOH + Pb(NO ₃) ₂	2 NaNO ₃ + Pb(OH) ₂
	Is this Balanced?	Yes

Balanced Reactions

	Reactant	Product
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{OH}$
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{OH}$

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{CO}_3$
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{S}$
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{OH}$
i	$\text{Na}_2\text{SO}_4 + \text{Pb}(\text{NO}_3)_2$	$(\text{NO}_3)_2 + \text{SO}_4$

Balanced Reactions

	Reactant	Product
f	$\text{Na}_2\text{CO}_3 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{CO}_3$
g	$\text{Na}_2\text{S} + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{S}$
h	$\text{NaOH} + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{OH}$
i	$\text{Na}_2\text{SO}_4 + \text{Pb}(\text{NO}_3)_3$	$(\text{NO}_3)_3 + \text{SO}_4$

Program of Studies

The background of the slide is a dark blue gradient. In the bottom right corner, there is a decorative graphic consisting of four parallel diagonal lines that slope downwards from left to right. The lines are a lighter shade of blue than the background.

Packet Points

	Points Possible
■ Section Review $9+7+7$	23
■ Chapter Review $10 + 7$	17
■ Activity (softdrink lab)	40
■ Notes/Vocabulary	20
■ TOTAL	100

Packet Points

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■ Section Review $9+7+7$	23
■ Chapter Review $10 + 7$	17
■ Activity (softdrink lab)	40
■ Notes/Vocabulary	20
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Packet Points

	Points Possible
■ Section Review $9+7+7$	23
■ Chapter Review $10 + 7$	17
■ Activity (softdrink lab)	40
■ Notes/Vocabulary	20
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